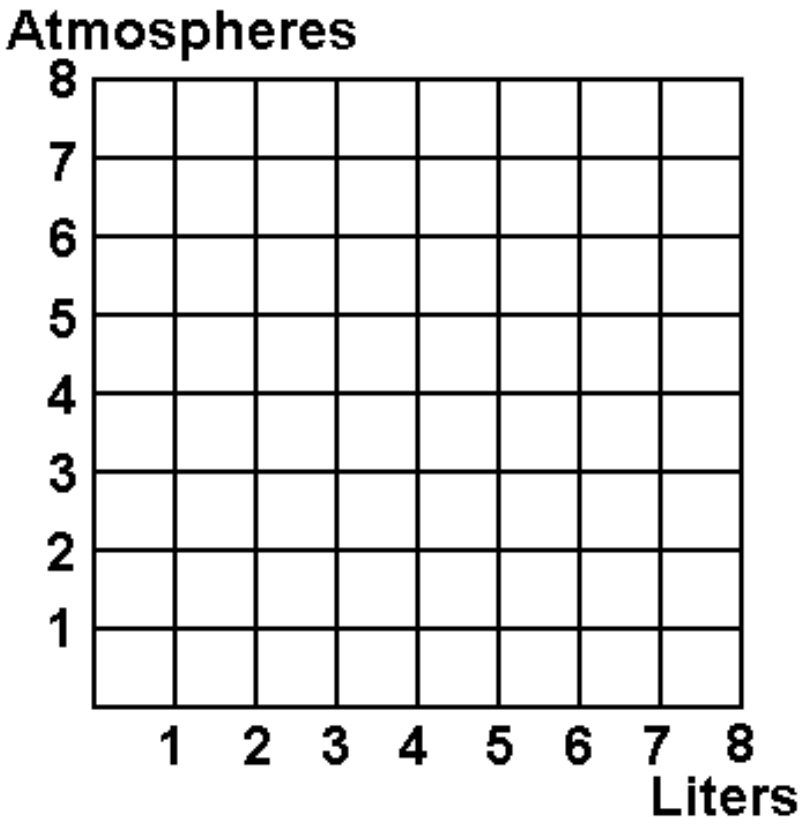
Handout 2 - Cyclic Process with Ideal Gas - Day 4



Given: $T_A = 100\text{ K} \xrightarrow{\text{isobaric}} T_B = 300\text{ K} \xrightarrow{\text{isovolumetric}} T_C = 100\text{ K} \xrightarrow{\text{isothermal}} T_A = 100\text{ K}$

$P_A = 4$ atm

$V_A = 1$ liter

Answer in terms of $nR = \dfrac{P_A V_A}{T_A} = \dfrac{(4\text{ atm})(1\text{ liter})}{100\text{ K}} = .04\ \dfrac{\text{lt}\cdot\text{atm}}{\text{K}}$

Find in terms of K, atm, and l:

	A	B	C
T			
P			
V			
U			

Determine in terms of K, atm, and l:

	A6B	B6C	C6A	A6B6C6A
Q				
W				
ÎU				